**Chapter 10 Vocabulary**

1. **Bond dipole**: This is due to unequal electron sharing between two atoms in a covalent bond.
2. **VSEPR Model**: This accounts for the geometry arrangements of shared and unshared electron pairs around a central atom in terms of the repulsions between electron pairs.
3. **Bonding pairs**: In a Lewis structure, these are shared by two atoms.
4. **Nonbonding pairs**: In a Lewis structure, these are assigned completely to one atom.
5. **Electron domains**: In the VSEPR model, regions about a central atom in which electrons are concentrated.
6. **Electron domain geometry** The three-dimensional arrangement of the electron domains around an atom, according to the VSEPR model.
7. **Molecular geometry**: The arrangement in space of the atoms of a molecule.
8. **Bond angles:** These are made by the lines joining the nuclei of the atoms in a molecule.